

**NEET(UG)-2018 TEST PAPER WITH ANSWER  
(HELD ON SUNDAY 06<sup>th</sup> MAY, 2018)**

**CHEMISTRY**

46. A mixture of 2.3 g formic acid and 4.5 g oxalic acid is treated with conc.  $\text{H}_2\text{SO}_4$ . The evolved gaseous mixture is passed through KOH pellets. Weight (in g) of the remaining product at STP will be  
(1) 1.4      (2) 3.0      (3) 2.8      (4) 4.4

**Ans. (3)**

47. Nitration of aniline in strong acidic medium also gives m-nitroaniline because  
(1) In spite of substituents nitro group always goes to only m-position.  
(2) In electrophilic substitution reactions amino group is meta directive.  
(3) In absence of substituents nitro group always goes to m-position  
(4) In acidic (strong) medium aniline is present as anilinium ion.

**Ans. (4)**

48. Which of the following oxides is most acidic in nature?  
(1) MgO      (2) BeO      (3) BaO      (4) CaO

**Ans. (2)**

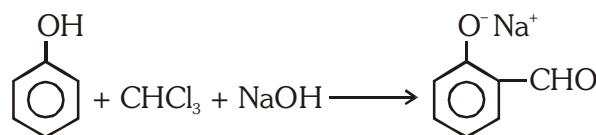
49. The difference between amylose and amylopectin is  
(1) Amylopectin have 1 → 4  $\alpha$ -linkage and 1 → 6  $\alpha$ -linkage  
(2) Amylose have 1 → 4  $\alpha$ -linkage and 1 → 6  $\beta$ -linkage  
(3) Amylopectin have 1 → 4  $\alpha$ -linkage and 1 → 6  $\beta$ -linkage  
(4) Amylose is made up of glucose and galactose

**Ans. (1)**

50. Regarding cross-linked or network polymers, which of the following statements is **incorrect**?  
(1) They contain covalent bonds between various linear polymer chains.  
(2) They are formed from bi- and tri-functional monomers.  
(3) Examples are bakelite and melamine.  
(4) They contain strong covalent bonds in their polymer chains.

**Ans. (4)**

51. In the reaction



the electrophile involved is

- (1) dichloromethyl cation ( $\text{CHCl}_2^+$ )  
(2) formyl cation ( $\text{CHO}^+$ )  
(3) dichloromethyl anion ( $\text{CHCl}_2^-$ )  
(4) dichlorocarbene ( $:\text{CCl}_2$ )

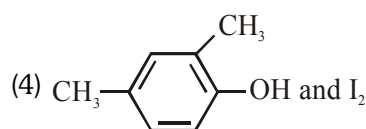
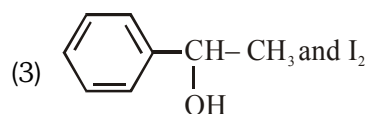
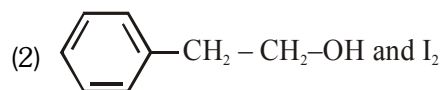
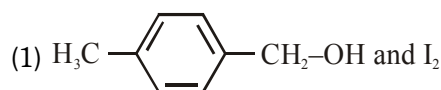
**Ans. (4)**

52. Carboxylic acid have higher boiling points than aldehydes, ketones and even alcohols of comparable molecular mass. It is due to their  
(1) formation of intramolecular H-bonding  
(2) formation of carboxylate ion  
(3) more extensive association of carboxylic acid via van der Waals force of attraction  
(4) formation of intermolecular H-bonding.

**Ans. (4)**

53. Compound A,  $\text{C}_8\text{H}_{10}\text{O}$ , is found to react with NaOI (produced by reacting Y with NaOH) and yields a yellow precipitate with characteristic smell.

A and Y are respectively



**Ans. (3)**

54. The correct difference between first- and second-order reaction is that

- (1) the rate of a first-order reaction does not depend on reactant concentration; the rate of a second-order reaction does depend on reactant concentrations.
- (2) the half-life of a first-order reaction does not depend on  $[A]_0$ ; the half-life of a second-order reaction does depend on  $[A]_0$
- (3) a first-order reaction can be catalyzed; a second-order reaction cannot be catalyzed.
- (4) the rate of a first-order reaction does depend on reactant concentrations; the rate of a second-order reaction does not depend on reactant concentrations

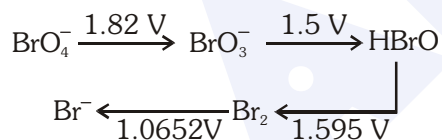
Ans. (2)

55. Among  $\text{CaH}_2$ ,  $\text{BeH}_2$ ,  $\text{BaH}_2$ , the order of ionic character is

- (1)  $\text{BeH}_2 < \text{CaH}_2 < \text{BaH}_2$
- (2)  $\text{CaH}_2 < \text{BeH}_2 < \text{BaH}_2$
- (3)  $\text{BeH}_2 < \text{BaH}_2 < \text{CaH}_2$
- (4)  $\text{BaH}_2 < \text{BeH}_2 < \text{CaH}_2$

Ans. (1)

56. Consider the change in oxidation state of Bromine corresponding to different emf values as shown in the diagram below:



Then the species undergoing disproportionation is:-

- (1)  $\text{BrO}_3^-$
- (2)  $\text{BrO}_4^-$
- (3)  $\text{Br}_2$
- (4)  $\text{HBrO}$

Ans. (4)

57. In which case is the number of molecules of water maximum?

- (1) 18 mL of water
- (2) 0.18 g of water
- (3) 0.00224 L of water vapours at 1 atm and 273 K
- (4)  $10^{-3}$  mol of water

Ans. (1)

58. Magnesium reacts with an element (X) to form an ionic compound. If the ground state electronic configuration of (X) is  $1s^2 2s^2 2p^3$ , the simplest formula for this compound is

- (1)  $\text{Mg}_2\text{X}_3$
- (2)  $\text{MgX}_2$
- (3)  $\text{Mg}_2\text{X}$
- (4)  $\text{Mg}_3\text{X}_2$

Ans. (4)

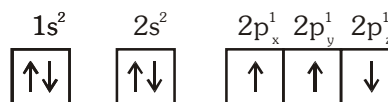
59. Iron exhibits bcc structure at room temperature. Above  $900^\circ\text{C}$ , it transforms to fcc structure. The ratio of density of iron at room temperature to that at  $900^\circ\text{C}$  (assuming molar mass and atomic radii of iron remains constant with temperature) is

- (1)  $\frac{\sqrt{3}}{\sqrt{2}}$
- (2)  $\frac{4\sqrt{3}}{3\sqrt{2}}$
- (3)  $\frac{3\sqrt{3}}{4\sqrt{2}}$
- (4)  $\frac{1}{2}$

Ans. (3)

60. Which one is a **wrong** statement ?

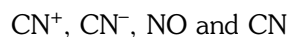
- (1) Total orbital angular momentum of electron in 's' orbital is equal to zero
- (2) An orbital is designated by three quantum numbers while an electron in an atom is designated by four quantum numbers.
- (3) The electronic configuration of N atom is



- (4) The value of m for  $d_{z^2}$  is zero

Ans. (3)

61. Consider the following species:



Which one of these will have the highest bond order?

- (1) NO
- (2)  $\text{CN}^-$
- (3)  $\text{CN}^+$
- (4) CN

Ans. (2)

62. Which of the following statements is **not** true for halogens ?

- (1) All form monobasic oxyacids.
- (2) All are oxidizing agents.
- (3) All but fluorine show positive oxidation states.
- (4) Chlorine has the highest electron-gain enthalpy.

Ans. (Bonus)

63. Which one of the following elements is unable to form  $MF_6^{3-}$  ion ?

- (1) Ga (2) Al (3) B (4) In

Ans. (3)

64. In the structure of  $ClF_3$ , the number of lone pairs of electrons on central atom 'Cl' is

- (1) one (2) two (3) four (4) three

Ans. (2)

65. Considering Ellingham diagram, which of the following metals can be used to reduce alumina ?

- (1) Fe (2) Zn (3) Mg (4) Cu

Ans. (3)

66. The correct order of atomic radii in group 13 elements is

- (1)  $B < Al < In < Ga < Tl$   
 (2)  $B < Al < Ga < In < Tl$   
 (3)  $B < Ga < Al < Tl < In$   
 (4)  $B < Ga < Al < In < Tl$

Ans. (4)

67. The correct order of N-compounds in its decreasing order of oxidation states is

- (1)  $HNO_3, NO, N_2, NH_4Cl$   
 (2)  $HNO_3, NO, NH_4Cl, N_2$   
 (3)  $HNO_3, NH_4Cl, NO, N_2$   
 (4)  $NH_4Cl, N_2, NO, HNO_3$

Ans. (1)

68. On which of the following properties does coagulating power of an ion depend ?

- (1) The magnitude of the charge on the alone  
 (2) Size of the ion alone  
 (3) Both magnitude and sign of the charge the ion  
 (4) The sign of charge on the ion alone

Ans. (3)

69. Following solutions were prepared by mixing different volumes of NaOH and HCl of different concentrations :

- a.  $60\text{mL } \frac{M}{10} \text{HCl} + 40\text{mL } \frac{M}{10} \text{NaOH}$   
 b.  $55\text{mL } \frac{M}{10} \text{HCl} + 45\text{mL } \frac{M}{10} \text{NaOH}$   
 c.  $75\text{mL } \frac{M}{5} \text{HCl} + 25\text{mL } \frac{M}{5} \text{NaOH}$   
 d.  $100\text{mL } \frac{M}{10} \text{HCl} + 100\text{mL } \frac{M}{10} \text{NaOH}$

pH of which one of them will be equal to 1 ?

- (1) b (2) a  
 (3) d (4) c

Ans. (4)

70. The solubility of  $BaSO_4$  in water  $2.42 \times 10^3 \text{ gL}^{-1}$  at 298 K. The value of solubility product ( $K_{sp}$ ) will be

(Given molar mass of  $BaSO_4 = 233 \text{ g mol}^{-1}$ )

- (1)  $1.08 \times 10^{-10} \text{ mol}^2 \text{ L}^{-2}$   
 (2)  $1.08 \times 10^{-12} \text{ mol}^2 \text{ L}^{-2}$   
 (3)  $1.08 \times 10^{-14} \text{ mol}^2 \text{ L}^{-2}$   
 (4)  $1.08 \times 10^{-8} \text{ mol}^2 \text{ L}^{-2}$

Ans. (1)

71. Given van der Waals constant for  $NH_3, H_2$  and  $CO_2$  are respectively 4.17, 0.244, 1.36 and 3.59, which one of the following gases is most easily liquefied?

- (1)  $NH_3$  (2)  $H_2$   
 (3)  $O_2$  (4)  $CO_2$

Ans. (1)

72. The compound A on treatment with Na gives B, and with  $PCl_5$  gives C. B and C react together to give diethyl ether. A, B and C are in the order

- (1)  $C_2H_5OH, C_2H_6, C_2H_5Cl$   
 (2)  $C_2H_5OH, C_2H_5Cl, C_2H_5ONa$   
 (3)  $C_2H_5Cl, C_2H_6, C_2H_5OH$   
 (4)  $C_2H_5OH, C_2H_5ONa, C_2H_5Cl$

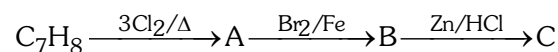
Ans. (4)

73. Hydrocarbon (A) reacts with bromine by substitution to form an alkyl bromide which by Wurtz reaction is converted to gaseous hydrocarbon containing less than four carbon atoms. (A) is

- (1)  $CH \equiv CH$  (2)  $CH_2 = CH_2$   
 (3)  $CH_3 - CH_3$  (4)  $CH_4$

Ans. (4)

74. The compound  $C_7H_8$  undergoes the following reactions :



The product 'C' is

- (1) m-bromotoluene  
 (2) o-bromotoluene  
 (3) 3-bromo-2,4,6-trichlorotoluene  
 (4) p-bromotoluene

Ans. (1)

75. Which oxide of nitrogen is **not** a common pollutant introduced into the atmosphere both due to natural and human activity ?

- (1)  $N_2O_5$
- (2)  $NO_2$
- (3)  $N_2O$
- (4)  $NO$

Ans. (1)

76. For the redox reaction

$MnO_4^- + C_2O_4^{2-} + H^+ \longrightarrow Mn^{2+} + CO_2 + H_2O$   
the correct coefficients of the reactants for the balanced equation are

- | $MnO_4^-$ | $C_2O_4^{2-}$ | $H^+$ |
|-----------|---------------|-------|
| (1) 16    | 5             | 2     |
| (2) 2     | 5             | 16    |
| (3) 2     | 16            | 5     |
| (4) 5     | 16            | 2     |

Ans. (2)

77. Which one of the following conditions will favour maximum formation of the product in the reaction,  
 $A_2(g) + B_2(g) \rightleftharpoons X_2(g) \quad \Delta_r H = -X \text{ kJ ?}$

- (1) Low temperature and high pressure
- (2) Low temperature and low pressure
- (3) High temperature and high pressure
- (4) High temperature and low pressure

Ans. (1)

78. The correction factor 'a' to the ideal gas equation corresponds to

- (1) density of the gas molecules
- (2) volume of the gas molecules
- (3) electric field present between the gas molecules
- (4) forces of attraction between the gas molecules

Ans. (4)

79. When initial concentration of the reactant is doubled, the half-life period of a zero order reaction

- (1) is halved
- (2) is doubled
- (3) is tripled
- (4) remains unchanged

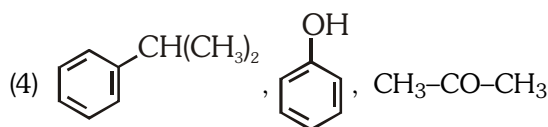
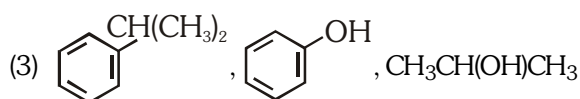
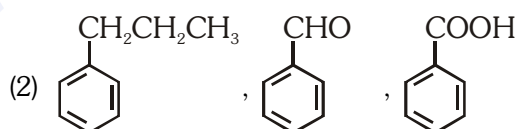
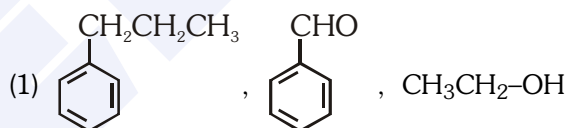
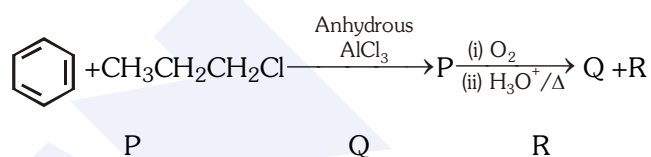
Ans. (2)

80. The bond dissociation energies of  $X_2$ ,  $Y_2$  and  $XY$  are in the ratio of 1 : 0.5 : 1.  $\Delta H$  for the formation of  $XY$  is  $-200 \text{ kJ mol}^{-1}$ . The bond dissociation energy of  $X_2$  will be

- (1)  $200 \text{ kJ mol}^{-1}$
- (2)  $100 \text{ kJ mol}^{-1}$
- (3)  $800 \text{ kJ mol}^{-1}$
- (4)  $400 \text{ kJ mol}^{-1}$

Ans. (3)

81. Identify the major products P, Q and R in the following sequence of reaction :



Ans. (4)

82. Which of the following compounds can form a zwitterion ?

- |                  |                 |
|------------------|-----------------|
| (1) Aniline      | (2) Acetanilide |
| (3) Benzoic acid | (4) Glycine     |

Ans. (4)

83. The type of isomerism shown by the complex  $[\text{CoCl}_2(\text{en})_2]$  is
- (1) Geometrical isomerism
  - (2) Coordination isomerism
  - (3) Ionization isomerism
  - (4) Linkage isomerism

Ans. (1)

84. Which one of the following ions exhibits d-d transition and paramagnetism as well ?

- (1)  $\text{CrO}_4^{2-}$
- (2)  $\text{Cr}_2\text{O}_7^{2-}$
- (3)  $\text{MnO}_4^-$
- (4)  $\text{MnO}_4^{2-}$

Ans. (4)

85. The geometry and magnetic behaviour of the complex  $[\text{Ni}(\text{CO})_4]$  are

- (1) square planar geometry and diamagnetic
- (2) tetrahedral geometry and diamagnetic
- (3) square planar geometry and paramagnetic
- (4) tetrahedral geometry and paramagnetic

Ans. (2)

86. Iron carbonyl,  $\text{Fe}(\text{CO})_5$  is

- (1) tetranuclear
- (2) mononuclear
- (3) trinuclear
- (4) dinuclear

Ans. (2)

87. Match the metal ions given in Column I with the spin magnetic moments of the ions given in Column II and assign the **correct** code :

Column I

Column II

- |                     |                      |
|---------------------|----------------------|
| a. $\text{Co}^{3+}$ | i. $\sqrt{8}$ B.M.   |
| b. $\text{Cr}^{3+}$ | ii. $\sqrt{35}$ B.M. |
| c. $\text{Fe}^{3+}$ | iii. $\sqrt{3}$ B.M. |
| d. $\text{Ni}^{2+}$ | iv. $\sqrt{24}$ B.M. |
|                     | v. $\sqrt{15}$ B.M.  |

- |     | a   | b  | c   | d   |
|-----|-----|----|-----|-----|
| (1) | iv  | v  | ii  | i   |
| (2) | i   | ii | iii | iv  |
| (3) | iv  | i  | ii  | iii |
| (4) | iii | v  | i   | ii  |

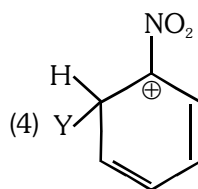
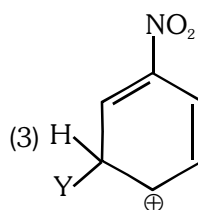
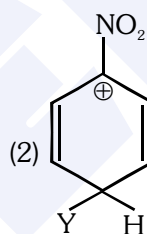
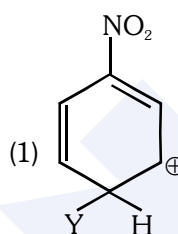
Ans. (1)

88. Which of the following is correct with respect to -I effect of the substituents ? (R = alkyl)

- (1)  $-\text{NH}_2 < -\text{OR} < -\text{F}$
- (2)  $-\text{NR}_2 < -\text{OR} < -\text{F}$
- (3)  $-\text{NH}_2 > -\text{OR} > -\text{F}$
- (4)  $-\text{NR}_2 > -\text{OR} > -\text{F}$

Ans. (1/2)

89. Which of the following carbocations is expected to be most stable ?



Ans. (3)

90. Which of the following molecules represents the order of hybridisation  $sp^2$ ,  $sp^2$ ,  $sp$ ,  $sp$  from left to right atoms ?

- (1)  $\text{HC} \equiv \text{C} - \text{C} \equiv \text{CH}$
- (2)  $\text{CH}_2 = \text{CH} - \text{C} \equiv \text{CH}$
- (3)  $\text{CH}_2 = \text{CH} - \text{CH} = \text{CH}_2$
- (4)  $\text{CH}_3 - \text{CH} = \text{CH} - \text{CH}_3$

Ans. (2)